**Liquids, Solids, Solutions and Intermolecular Forces Study Guide**

**Chapters 11-12**

**Solids**

* Crystalline solid
* Amorphous solid
* Lattice Energy

**Intermolecular Forces**

* London Dispersion Force
* Dipole-Dipole force
* Hydrogen bonding
* Ion-Dipole Force
* Surface tension
* Viscosity
* Capillary Action
* Volatile
* Chromatography

**Phase Changes**

* ΔHvaporization
* ΔHfusion
* Vaporization
* Condensation
* Melting
* Freezing
* Sublimation
* Deposition
* Phase Diagrams
* Triple Point
* Critical Point
* Dynamic Equilibrium
* Vapor Pressure
* Boiling Point
* Normal Boiling Point
* Melting point

**Solutions**

* Solution
* Solute
* Solvent
* Aqueous
* Solubility
* Entropy
* Dilute
* Concentratied
* Saturated
* Unsaturated
* Supersaturated
* Recrystallization
* Temperature effects on solubility of Solids and Gases in water
* Molarity
* Molarity by dilution (M1V1=M2V2)
* Miscibility (“like dissolves like”)

**Suggested Figures to Review**

* 11.2
* 11.3
* Table 11.3
* 11.6
* 11.11
* Table 11.5
* Table 11.6
* 11.22
* 11.25
* 11.27
* 11.28
* 11.29
* 11.35
* 11.36
* 11.37
* 12.4
* Table 12.2
* 12.8
* 12.9
* 12.11
* 12.12
* 12.13